

Centre Number	Candidate Number	Name
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UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CHEMISTRY

0620/03

Paper 3

May/June 2004

1 hour 15 minutes

Candidates answer on the Question Paper.
No Additional Materials required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.
Write in dark blue or black pen in the spaces provided on the Question Paper.
You may use a pencil for any diagrams, graphs or rough working.
Do not use staples, paper clips, highlighters, glue or correction fluid.
You may use a calculator.

Answer **all** questions.
The number of marks is given in brackets [] at the end of each question or part question.
A copy of the Periodic Table is printed on page 12.

For Examiner's Use	
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2	
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Total	

If you have been given a label, look at the details. If any details are incorrect or missing, please fill in your correct details in the space given at the top of this page.

Stick your personal label here, if provided.

This document consists of **12** printed pages.

1 It was reported from America that a turbine engine, the size of a button, might run on batteries. The engine would be built from silicon which has suitable properties for this purpose.

(a) (i) Why are batteries a convenient source of energy?

..... [1]

(ii) The engine will run on a small pack of jet fuel. What other chemical is needed to burn this fuel?

..... [1]

(b) Silicon has the same type of macromolecular structure as diamond.

(i) Explain why one atom of either element can form four covalent bonds.

.....
..... [2]

(ii) Predict **two** physical properties of silicon.

.....
..... [2]

(iii) Name a different element that has a similar structure and properties to silicon.

..... [1]

(c) Silicon is made by the carbon reduction of the macromolecular compound, silicon(IV) oxide.

(i) Balance the equation for the reduction of silicon(IV) oxide.



(ii) Explain why the silicon(IV) oxide is said to be reduced.

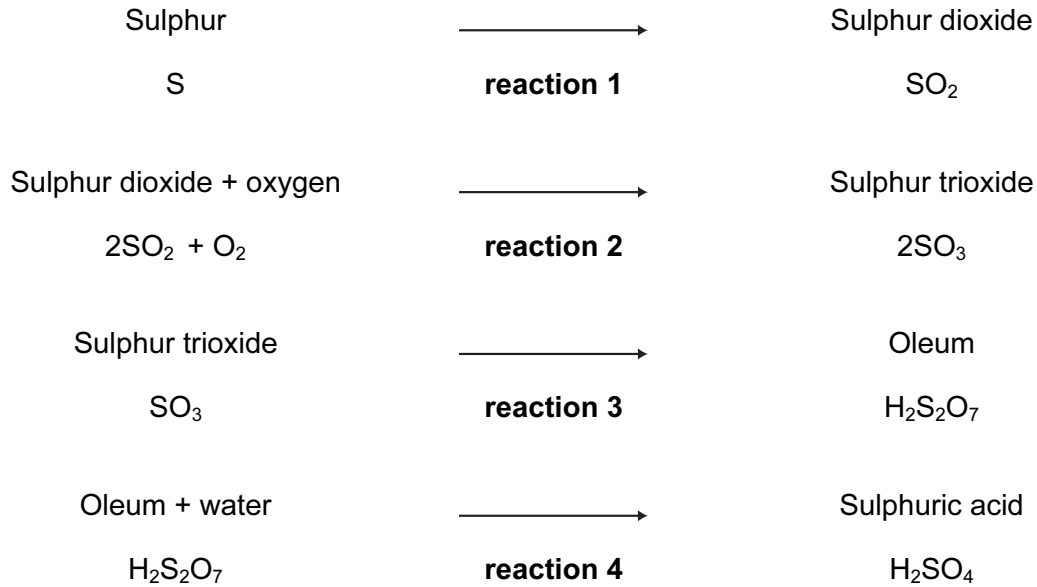
..... [1]

(iii) Describe the structure of silicon(IV) oxide. You may use a diagram.

..... [2]

2 Sulphur is used to make sulphuric acid. In the UK, the annual production of the acid is about 2.5 million tonnes.

(a) The reactions in the manufacture of sulphuric acid by the Contact Process are shown below.



(i) Give a large scale source of the element sulphur.

..... [1]

(ii) State another use of sulphur dioxide.

..... [1]

(iii) How is sulphur changed into sulphur dioxide?

..... [1]

(iv) Name the catalyst used in reaction 2.

..... [1]

(v) Reaction 2 is exothermic. Why is a catalyst, rather than a higher temperature, used to increase the rate of this reversible reaction?

.....
..... [2]

(vi) Write a word equation for reaction 3.

..... [1]

(vii) Write a symbol equation for reaction 4.

..... [1]

(b) About one third of this production of acid is used to make nitrogen and phosphorus-containing fertilisers.

(i) Name the third element that is essential for plant growth and is present in most fertilisers.

..... [1]

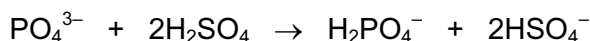
(ii) Name a nitrogen-containing fertiliser that is manufactured from sulphuric acid.

..... [1]

(iii) Rock phosphate (calcium phosphate) is obtained by mining. It reacts with concentrated sulphuric acid to form the fertiliser, superphosphate. Predict the formula of each of these phosphates.

fertiliser	ions	formula
calcium phosphate	Ca^{2+} and PO_4^{3-}
calcium superphosphate	Ca^{2+} and H_2PO_4^- [2]

(iv) The ionic equation for the reaction between the phosphate ion and sulphuric acid is shown below.



Explain why the phosphate ion is described as acting as a base in this reaction.

..... [2]

3 An organic compound decomposes to form nitrogen.



(a) Explain the state symbols.

aq

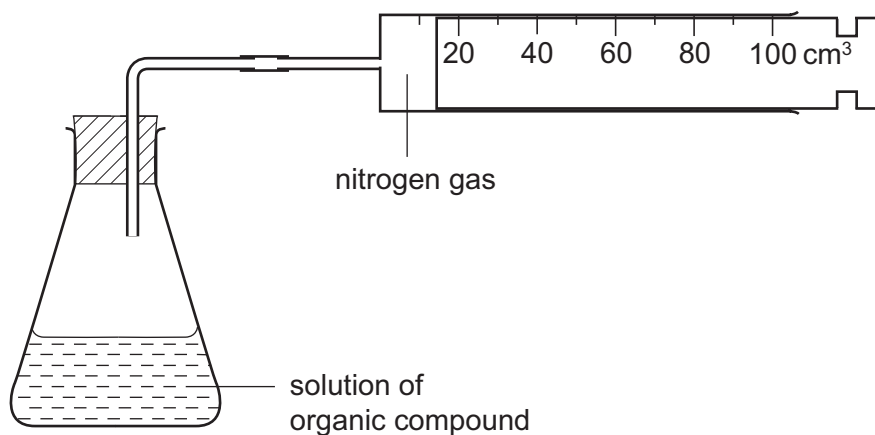
l

g [2]

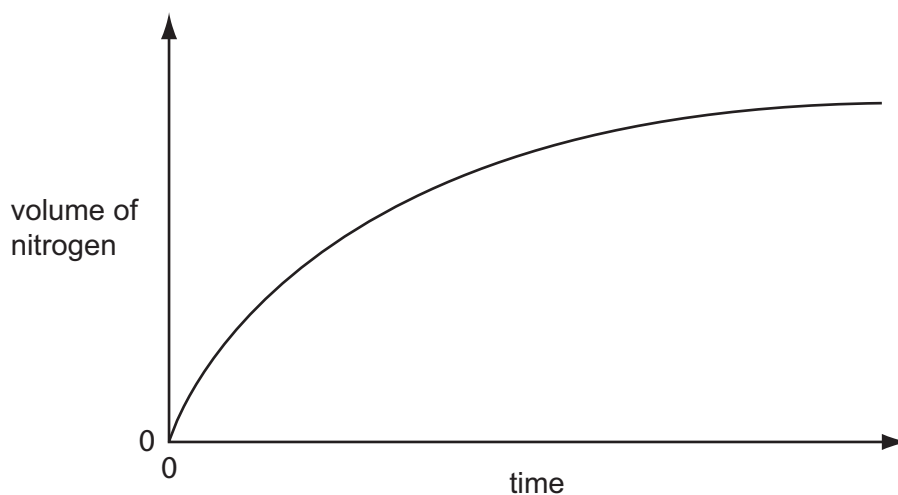
(b) Draw a diagram to show the arrangement of the valency electrons in **one** molecule of nitrogen.

[2]

(c) The rate of this reaction can be measured using the following apparatus.



The results of this experiment are shown on the graph below.



(i) How does the rate of this reaction vary with time?

.....
 [1]

(ii) Why does the rate vary?

.....
 [2]

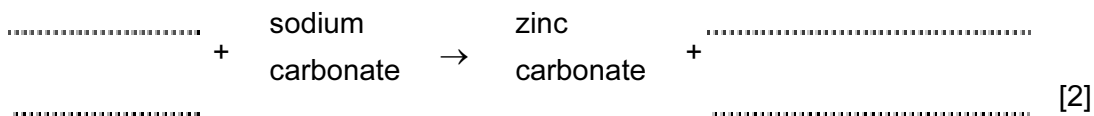
(iii) The reaction is catalysed by copper powder. Sketch the graph for the catalysed reaction on the same grid. [2]

(iv) Why is copper powder more effective as a catalyst than a single piece of copper?

..... [1]

4 (a) Insoluble compounds are made by precipitation.

(i) Complete the word equation for the preparation of zinc carbonate.



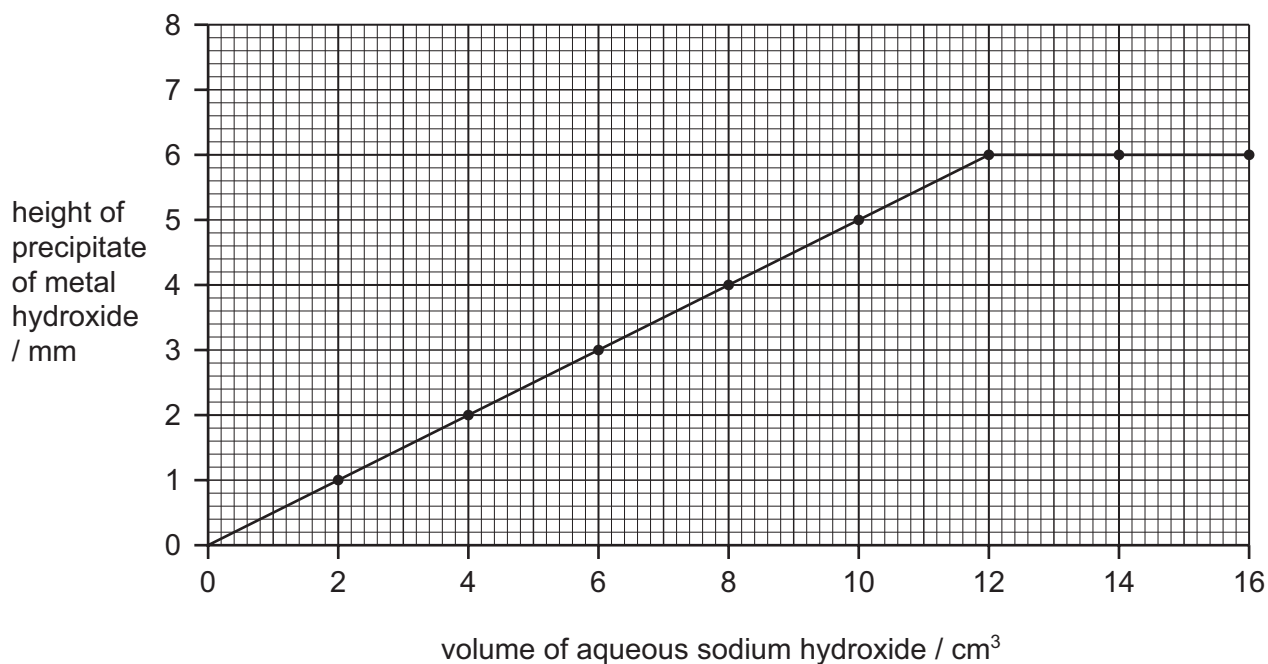
(ii) Complete the following symbol equation.



(iii) Write an ionic equation for the precipitation of the insoluble salt, silver(I) chloride.



(b) 2.0 cm^3 portions of aqueous sodium hydroxide were added to 4.0 cm^3 of aqueous iron(III) chloride. Both solutions had a concentration of 1.0 mol/dm^3 . After each addition, the mixture was stirred, centrifuged and the height of the precipitate of iron(III) hydroxide was measured. The results are shown on the following graph.



(i) Complete the ionic equation for the reaction.



(ii) On the same grid, sketch the graph that would have been obtained if iron(II) chloride had been used instead of iron(III) chloride? [2]

- (iii) If aluminium chloride had been used instead of iron(III) chloride, the shape of the graph would be different. How are the shapes of these two graphs different and why?

difference in shape

.....

reason for difference

..... [2]

- 5 (a) Copper has the structure of a typical metal. It has a lattice of positive ions and a "sea" of mobile electrons. The lattice can accommodate ions of a different metal.

Give a **different** use of copper that depends on each of the following.

- (i) the ability of the ions in the lattice to move past each other

..... [1]

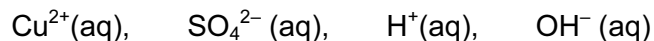
- (ii) the presence of mobile electrons

..... [1]

- (iii) the ability to accommodate ions of a different metal in the lattice

..... [1]

- (b) Aqueous copper(II) sulphate solution can be electrolysed using carbon electrodes. The ions present in the solution are as follows.



- (i) Write an ionic equation for the reaction at the negative electrode (cathode).

..... [1]

- (ii) A colourless gas was given off at the positive electrode (anode) and the solution changes from blue to colourless.

Explain these observations.

.....

..... [2]

- (c) Aqueous copper(II) sulphate can be electrolysed using copper electrodes. The reaction at the negative electrode is the same but the positive electrode becomes smaller and the solution remains blue.

(i) Write a word equation for the reaction at the positive electrode.

..... [1]

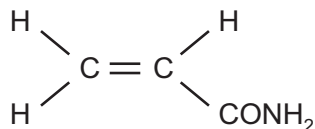
(ii) Explain why the colour of the solution does not change.

.....
..... [2]

(iii) What is the large scale use of this electrolysis?

..... [1]

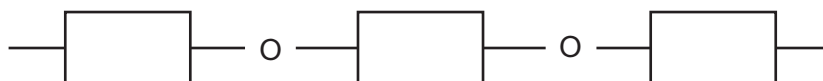
- 6 In 2002, Swedish scientists found high levels of acrylamide in starchy foods that had been cooked above 120 °C. Acrylamide, which is thought to be a risk to human health, has the following structure.



- (a) (i) It readily polymerises to polyacrylamide. Draw the structure of this polymer.

[2]

- (ii) Starch is formed by polymerisation. It has a structure of the type shown below. Name the monomer.



[1]

- (iii) What are the differences between these two polymerisation reactions, one forming polyacrylamide and the other starch?

[2]

- (b) Acrylamide hydrolyses to form acrylic acid and ammonium ions.

- (i) Describe the test for the ammonium ion.

test

.....

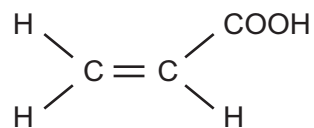
result

..... [2]

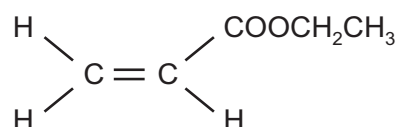
- (ii) Given an aqueous solution, concentration 0.1 mol / dm³, how could you show that acrylic acid is a weak acid.

.....
 [2]

- (c) The structural formula of acrylic acid is shown below. It forms compounds acrylates.



- (i) Acrylic acid reacts with ethanol to form the following compound.



Deduce the name of this compound. What type of organic compound is it?

name

type of compound [2]

- (ii) Acrylic acid is an unsaturated compound. It will react with bromine. Describe the colour change and draw the structural formula of the product of this addition reaction.

colour change

structural formula of product

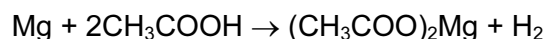
[2]

- 7 Chemists use the concept of the mole to calculate the amounts of chemicals involved in a chemical reaction.

(a) Define *mole*.

..... [1]

(b) 3.0 g of magnesium was added to 12.0 g of ethanoic acid.



The mass of one mole of Mg is 24 g.

The mass of one mole of CH₃COOH is 60 g.

(i) Which one, magnesium or ethanoic acid, is in excess? You must show your reasoning.

..... [3]

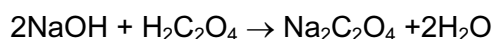
(ii) How many moles of hydrogen were formed?

..... [1]

(iii) Calculate the volume of hydrogen formed, measured at r.t.p.

..... [2]

(c) In an experiment, 25.0 cm³ of aqueous sodium hydroxide, 0.4 mol/dm³, was neutralised by 20.0 cm³ of aqueous oxalic acid, H₂C₂O₄.



Calculate the concentration of the oxalic acid in mol/dm³.

(i) Calculate the number of moles of NaOH in 25.0 cm³ of 0.4 mol/dm³ solution.

..... [1]

(ii) Use your answer to (i) and the mole ratio in the equation to find out the number of moles of H₂C₂O₄ in 20 cm³ of solution.

..... [1]

(iii) Calculate the concentration, mol/dm³, of the aqueous oxalic acid.

..... [2]

DATA SHEET
The Periodic Table of the Elements

		Group																	
		I	II	III	IV	V	VI	VII	VIII	IX	X								
7	9	Li Lithium 3	Be Beryllium 4	<table border="1" style="margin: auto; border-collapse: collapse;"> <tr> <td style="text-align: center;">1</td> <td style="text-align: center;">H Hydrogen 1</td> </tr> </table>										1	H Hydrogen 1				
1	H Hydrogen 1																		
23	24	Na Sodium 11	Mg Magnesium 12																
39	40	K Potassium 19	Ca Calcium 20	45	48	51	52	55	56	59	64	65	70	73	75	79	80	84	86
85	88	Rb Rubidium 37	Sr Strontium 38	Sc Scandium 21	Ti Titanium 22	V Vanadium 23	Cr Chromium 24	Mn Manganese 25	Fe Iron 26	Co Cobalt 27	Ni Nickel 28	Cu Copper 29	Zn Zinc 30	Ga Gallium 31	Ge Germanium 32	As Arsenic 33	Se Selenium 34	Br Bromine 35	Kr Krypton 36
133	137	Cs Caesium 55	Ba Barium 56	89	91	93	96	101	106	108	112	115	119	122	127	128	131	131	131
226	227	Fr Francium 87	Ra Radium 88	Y Yttrium 39	Zr Zirconium 40	Nb Niobium 41	Mo Molybdenum 42	Tc Technetium 43	Ru Ruthenium 44	Rh Rhodium 45	Pd Palladium 46	Ag Silver 47	Cd Cadmium 48	In Indium 49	Sn Tin 50	Sb Antimony 51	Te Tellurium 52	I Iodine 53	Xe Xenon 54
227	227	Ac Actinium 89	La Lanthanum 57	144	141	140	141	144	150	152	157	159	162	165	167	169	173	175	175
227	227	Fr Francium 87	Ra Radium 88	Pr Praseodymium 59	Pm Promethium 61	Ce Cerium 58	Pr Praseodymium 59	Nd Neodymium 60	Pm Promethium 61	Sm Samarium 62	Eu Europium 63	Gd Gadolinium 64	Tb Terbium 65	Dy Dysprosium 66	Ho Holmium 67	Er Erbium 68	Tm Thulium 69	Yb Ytterbium 70	Lu Lutetium 71
227	227	Fr Francium 87	Ra Radium 88	Pa Protactinium 91	Np Neptunium 93	Th Thorium 90	Pa Protactinium 91	U Uranium 92	Pu Plutonium 94	Am Americium 95	Cm Curium 96	Bk Berkelium 97	Cf Californium 98	Es Einsteinium 99	Fm Fermium 100	Md Mendelevium 101	No Nobelium 102	Lr Lawrencium 103	Lr Lawrencium 103

*58-71 Lanthanoid series
90-103 Actinoid series

Key

a	X	b
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a = relative atomic mass
X = atomic symbol
b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).